

Synthesis Tetrathioarsenate as a Precipitant of Ammoniate Ions of Transitional Metals

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Abstract: Synthesized substances obtained by sodium tetrathioarsenate reaction with silver(I), cobalt (II), mickelous (II), copper (II), zincous, cadmium and mercury have been studied by IR- spectroscopy, X-ray analysis.

Keywords: Tetrathioarsenate, Transitional Metals

1. Introduction

Absolute majority of transition metals amiacates are water-soluble complexes. Therefore their separation in individual state needs concentration by evaporation till obtaining saturated solution or re-precipitation with some solvent, that is rather difficult [1].

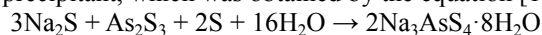
On the other hand, synthesis conditions of transition metals tetraoxoarsenates and their physical-chemical properties are rather well investigated [2], but data about tetrathioarsenates are limited enough [3-5]. As for d-metals application of tetrathioarsenates for obtaining corresponding coordinative compounds, we don't have any data about them.

The goal of the present work is to solve the following problems: 1) to establish the possibility of using tetrathioarsenates of alkaline metals as precipitants of ammoniate ions of transitional metals form aqueous solutions; 2) in case if the experiment is a success to study the products of reaction by chemical and physico-chemical methods and in this way to establish the mechanism of characteristic changes in new synthesized coordinative compounds in relation with the change of central atom.

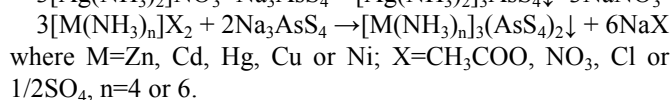
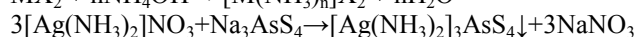
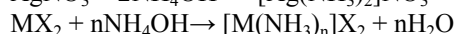
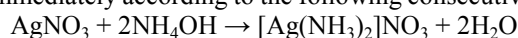
2. Research Methods

Silver (I) and mercury (II) nitrates, nickel (II) and cobalt (II) chlorides, zinc and cadmium acetates and copper (II) sulphide

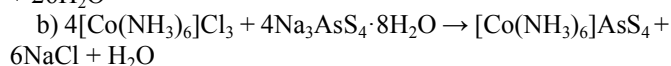
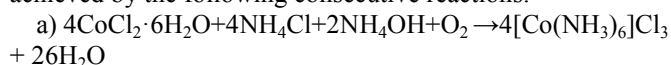
were used as initial substances and sodium tetrathioarsenate as a precipitant, which was obtained by the equation [1]:



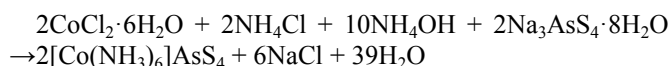
The experiment was carried out in the following way: first the ammoniates of transitional (d-elements) metals were obtained by the action of ammonium alkali in the appropriate water-soluble salts, then the reaction product (without isolating in the individual state) was treated by sodium tetrathioarsenate solution in the same aqueous solution. Tetrathioarsenate ammoniates of d-metals were precipitated immediately according to the following consecutive reactions:



All attempts to obtain tetrathioarsenate ammoniate of cobalt (II) in analogous way were unsuccessful. Cobalt (III) is found to be easily oxidized by atmospheric oxygen in ammonium alkaline solution. Therefore, we decided to make use of this fact to obtain tetrathioarsenate of cobalt (III). It can be achieved by the following consecutive reactions:



totally



Elementary analysis of synthesized substances was carried out by the well-known methods: arsenic was defined by the method of Ewins [2], sulphur by gravimetric.

Load of initial substances and yield of obtained products are presented in Table 1 and the results of chemical analysis of synthesized substances in Table 2.

Table 1. Load of Starting substances and yield of obtained products

	Load of starting materials						yield of obtained products		
	salt			ammonia	$\text{Na}_3\text{AsS}_4 \cdot 8\text{H}_2\text{O}$				
	formula	g	mole	solution (ml)	g	mole	g	mole	%
I	AgNO_3	1.84	0.0108	15.0	3.0	0.0036	2.27	0.0036	86.8
II	$\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$	2.70	0.0108	20.0	3.0	0.0072	2.72	0.0034	94.4
III	$\text{Zn}(\text{CH}_3\text{COO})_2 \cdot 2\text{H}_2\text{O}$	2.37	0.0108	16.0	3.0	0.0072	2.48	0.0031	85.8
IV	$\text{Cd}(\text{CH}_3\text{COO})_2 \cdot 2\text{H}_2\text{O}$	2.50	0.0107	16.0	3.0	0.0072	3.10	0.0033	90.9
V	$\text{Hg}(\text{NO}_3)_2 \cdot \text{H}_2\text{O}$	3.42	0.0108	20.0	3.0	0.0072	4.03	0.0033	92.8
VI	$\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$	2.57	0.0108	18.0	3.0	0.0072	2.84	0.0032	88.5
VII	$\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$	2.58	0.0108	18.0	3.0	0.0072	3.16	0.0035	96.3

Table 2. The results of chemical analysis of synthesized compounds

Compound	Colour	Found, %				Calculated, %			
		M	As	N	S	M	As	N	S
I	$[\text{Ag}(\text{NH}_3)_2]_3\text{AsS}_4$	51.38	11.84	13.48	20.47	51.51	11.92	13.96	20.35
II	$[\text{Cu}(\text{NH}_3)_4]_3(\text{AsS}_4)_2$	23.88	18.56	21.07	31.87	23.79	18.74	20.48	31.98
III	$[\text{Zn}(\text{NH}_3)_4]_3(\text{AsS}_4)_2$	24.76	18.54	20.97	31.44	24.34	18.61	20.84	31.76
IV	$[\text{Hg}(\text{NH}_3)_4]_3(\text{AsS}_4)_2$	49.52	12.23	13.92	21.24	49.66	12.98	13.86	21.13
V	$[\text{Cd}(\text{NH}_3)_4]_3(\text{AsS}_4)_2$	35.48	15.72	17.86	27.18	35.60	15.84	17.74	27.03
VI	$[\text{Ni}(\text{NH}_3)_6]_3(\text{AsS}_4)_2$	19.98	16.76	28.46	28.74	19.83	16.89	28.38	28.83
VII	$[\text{Co}(\text{NH}_3)_6]_3(\text{AsS}_4)_2$	19.82	16.74	28.50	28.96	19.61	16.87	28.35	28.80

3. Main Results

3.1. Constitution and Structure of Synthesized Substances

Synthesized complexes are fine crystalline substances, which are practically insoluble on water or any other organic solvents. They have not a definite melting point and by heating above 150°C they are decomposed. They are insoluble in alkalis except $[\text{Zn}(\text{NH}_3)_4]_3(\text{AsS}_4)_2$. Their reaction with acids (HCl , H_2SO_4) is a complex process and requires to be studied separately. We can make a foregone conclusion that one of the products of reaction is arsenic (V) sulphide. Their reaction with concentrated nitrous acid is the exception when As_4S_{10} is changed.

Constitution and structure of synthesized substances is confirmed by the data of IR-spectroscopy and roentgen-phase studies, apart from the chemical analysis.

Study of IR-spectra of the substances under consideration shows that in every sample there are noted bands of 470 cm^{-1} deformative vibration characteristic to $-\text{As}-\text{S}$ bond [4] and the bands of valency vibration are noted at 430 cm^{-1} regions [6]. Bands at 1610 cm^{-1} and 3150 cm^{-1} region belonging to modified deformative and valency vibrations of coordinated ammonium respectively, allow us to conclude that the obtained compounds are ammoniates of d-metals [7].

Individuality of products is conformed by roentgen-phase analysis as well as ---- relation with the change of cation. It was shown by various roentgen reflexions and their distribution.

3.2. Synthesis of $[\text{Ag}(\text{NH}_3)_2]_3\text{AsS}_4$

Saturated solution of 3.0 g (0.0176 mole) silver (I) nitrate was added by excess amount of alkyl ammonium until the first formed precipitation was dissolved. The solution obtained was treated with 2.45 g (0.0059 mole) of saturated solution of sodium tetrathioarsenate with constant stirring. Some black substance was precipitated at once. It was hold in the mother solution to be formed into crystal substance. Next day the flask contents was filtered and washed with water and alcohol. After drying it under the air 3.20 g (0.0051 mole) of $[\text{Ag}(\text{NH}_3)_2]_3\text{AsS}_4$ was obtained i.e. 86.4% of the theoretical.

Other ammoniate complexes were also obtained analogously except $[\text{Co}(\text{NH}_3)_6]\text{AsS}_4$.

3.3. Synthesis of $[\text{Co}(\text{NH}_3)_6]\text{AsS}_4$

Mixture of $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$ (3.0 g) and NH_4Cl (2.0 g) was dissolved in water (25 ml). The solution was added by 0.1 g of activated carbon, 15.0 ml of concentrated ammonium alkali and a strong stream of air was let into it until the solution changed from red into yellowish-brown. The flask content was filtered and added by saturated solution of 2.0 g (0.0048 mole) of $\text{Na}_3\text{AsS}_4 \cdot 8\text{H}_2\text{O}$ with constant stirring. The small-crystal substance of dark orange color precipitated at once. Next day it was filtered, washed with the diluted solution of ammonia and dried out in vacuum desiccators on water-free alkaline kalium until a permanent mass was obtained. In the result 2.0 g (0.0042 mole) of $[\text{Co}(\text{NH}_3)_6]\text{AsS}_4$ was obtained, i.e. 98.5% of the theoretical.

4. Conclusion

Thus, studies have shown that by treatment of silver (I), mercury (II), zinc, cadmium, copper (II), nickel (II) and cobalt (III) amiacates with sodium tetrathioarsenates without their separation in individual condition, exchange reaction takes place by formation corresponding complexes. Obtained complex compounds are cationic-complexes, and tetrathioarsenat-ion is located in the outer sphere.

There are certain regularity by thermal decomposition of obtained compounds. That is reflected above all in extracting of ligand – ammonia, following dethionating and removing arsenic in sulfide form (realgar).

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